Tying it all Together: Human Kinetics and Kinetic Molecular Theory

Today in class we spent a lot of time having fun outside learning about Kinetic Molecular Theory, but how do all of our “human examples” connect to KMT?

Fill in the chart below with our class’s representation of each postulate (assumption) of the Kinetic Molecular Theory. If you don’t remember what we did, make up another example from real life!

|  |  |
| --- | --- |
| Postulate | Our Example |
| ***Gases consist of large numbers of tiny particles that are far apart relative to their own size.*** |  |
| ***There are no forces of attraction or repulsion between gas particles.*** |  |
| ***Gas particles move continuously, rapidly, and randomly in straight lines in all directions.*** |  |
| ***All collisions between particles and each other or the container are considered to be elastic collisions (no loss of kinetic energy)*** |  |
| ***The average kinetic energy depends on the temperature of the gas (directly increases with temperature increase)*** |  |

The hypothetical gas described by the Kinetic Molecular Theory is called an \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_!